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Theoretical yield = 5.95g  $\text{KClO}_4 \times 1 \text{ mol} \times 122.55 \text{ g} \times 2 \text{ O}_2 \times 1 \text{ mol} \times 32.00 \text{ g} = 3.11 \text{ g O}_2$  %Yield =  $\frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$  34.  
 $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$  Percent purity = 62.8%? pure  $\text{FeCO}_3$  2th, 2024

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