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Practice Test Ch3 Stoichiometry (page 3 Of 3)  
1. D It Might Be Easiest To Balance The Equation  
With Mostly Whole Numbers:  $2 \text{NH}_3 + \frac{7}{2} \text{O}_2 \rightarrow 2 \text{NO}_2 + 3 \text{H}_2\text{O}$ .  
The Question Asks About The Amount Of Oxygen  
Reacting With ONE Mole Of Ammonia, Thus Cut The  $\frac{7}{2}$  (3.5) Of Oxygen  
In Half To 1.75  
2. C  
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