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26. According To The Kinetic-molecular Theory, The Particles In A Liquid

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Negative Charge Near The Oxygen End Of The Molecule And Partial Positive Charge
Near The Hydrogen Atoms. Thus, The Negative End Of A Water Molecule Is Strongly
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EXAMPLE PROBLEM: Determine The Molar Mass Of A Compound Calculate The Molar
Mass For Each Of The Following Compounds: (a) 2-Propanol, $\text{CH}_3\text{CH}(\text{OH})\text{CH}_3$ (b)
Iron(II) Phosphate, $\text{Fe}_3(\text{PO}_4)_2$ SOLUTION: You Are Asked To Calculate The Molar
Mass Of A Compound. You Are Given The Compound's Formula. May 9th, 2024.
CHAPTER 2: Atoms, Molecules And Stoichiometry 2.5 Stoichiometry And Equations

Stoichiometry 1) Stoichiometry Is The Proportion Of Things Either Reacting Or Combining. 2) In Compounds, It Refers To The Ratio In Which The Atoms Are Combined Together. For Example, Water, H_2O Has A Stoichiometry Of 2 Hydrogen To 1 Oxygen. 3) It Also Refers To The Reacting Proportions In A Chemical Equation.
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