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18 - 1 Chapter 18: Electrochemistry Oxidation States An Oxidation-reduction Reaction, Or Redox Reaction, Is One In Which Electrons Are Transferred.  $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$  Each Sodium Atom Is Losing One Electron To Form  $\text{Na}^+$   $\text{Na} \rightarrow \text{Na}^+ + 1\text{e}^-$  This Loss Of Electrons Is Called Oxidation. Each Chlorine Atom Is Gaining 1 Electron To Form  $\text{Cl}^-$   $\text{Cl}_2 + 2\text{e}^-$  Jan 1th, 2024

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Objective 9: Use The Nernst Equation To Calculate The Equilibrium Constant, K. Do Problems 13 And 14 At The End Of This Section. Do The Following End-of-chapter Problems: 72, 74, 78 Apr 1th, 2024

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