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Name: ___KEY___ 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1 \text{ mol}$
 $122.55 \text{ g} \times 2 \text{ O}_2$ $1 \text{ KClO}_4 \times 32.00 \text{ g} \text{ 1 mol} = 3.11 \text{ g O}_2$
 $\% \text{ Yield} = \frac{1.45 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$
 $3 \text{ FeCO}_3 + \text{O}_2 \rightarrow 2 \text{ Fe}_2\text{O}_3 + 4 \text{ CO}_2$
Percent purity = 62.8%? pure FeCO_3 13th, 2024

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