

# Chapter Test B Stoichiometry

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## **Practice Test Ch 3 Stoichiometry Name Per**

Practice Test Ch3 Stoichiometry (page 3 Of 3) 1. D It  
Might Be Easiest To Balance The Equation With Mostly  
Whole Numbers:  $2 \text{NH}_3 + \frac{7}{2} \text{O}_2 \rightarrow 2 \text{NO}_2 + 3 \text{H}_2\text{O}$ .  
The Question Asks About The Amount Of Oxygen  
Reacting With ONE Mole Of Ammonia, Thus Cut The  $\frac{7}{2}$   
(3.5) Of Oxygen In Half To 1.75 2. C Balance: CH Feb  
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Unit 10 Stoichiometry Test Answers Monday, February 13  
Agenda: 1) Go Over Quest 2) Introduce Stoichiometry-S'more Activity HW: S'More Worksheet #2  
Tuesday, February 14  
Agenda: Lab -Introduction To Stoichiometry HW: 1) Lab Calculations 2) Read P 237-240  
Wednesday, February 15  
Agenda: 1) Go Over Lab 2) Worksheet #3 HW: 1) Read P 242 2) Video Notes  
Thursday, Mar 3th, 2024

## **Stoichiometry Test Review March 21/22**

5. Given The Incomplete Equation:  $2 \text{N}_2\text{O}_5(\text{g}) \rightleftharpoons$  Which

Set Of Products Completes And Balances The Incomplete Equation? 6. Base Your Answer To The Following Question On Given The Balanced Equation Representing A Reaction:  $C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$  What Is The Total Number Of Moles Of  $O_2(g)$  May 11th, 2024

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