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=  $k[D]^1 X[B]^1 Y$   $k[D]^2 X[B]^2 Y$  Equation 5 In This Equation K Cancels Out.  $[D]^1 = [D]^2$  Because We Are Starting Off With The Same Initial Concentration Of A In Each Trial. The Order Of Species D Which Is X Is Also Constant. Canceling Terms We Have Left:  $Rate_1 = Rate_2 = k[B]^1 Y = k[B]^2 Y$  Equation 6 Consider The Following Chemical Reaction: May 3th, 2024

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2 and  $\text{Cl}_2$  to form  $\text{HCl}$ . 7 Chemistry-I. Figure 1.1: Plot Of Concentration Against T For A Reaction  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$  The Rate Of This Reaction Does Not Depend On Concentration. 1.6 First Order Kinetics A Reaction Of The first Order Is Represented As  $X \rightarrow Y$  Where X Is The Reactant And Y The Product. The Rate Of The Reaction Will Be Directly Jan 9th, 2024

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