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Chemistry Worksheet On Stoichiometry Mixed Review Answers Calcium  $\text{C}_2\text{O}_2 + 3 \text{Cr}(3)$  When The Potassium Chlorate ( $\text{KClO}_3$ ) Is Heated, It Decomposes To Form Gases  $\text{KCl}$  And Oxygen ( $\text{O}_2$ .)  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$  C. Jun 1th, 2024

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Stoichiometry Worksheet #1 Answers 1. Given The Following Equation:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ , Show What The Following Molar Ratios Should Be. A.  $\text{C}_4\text{H}_{10} / \text{O}_2$  B.  $\text{O}_2 / \text{CO}_2$  C.  $\text{O}_2 / \text{H}_2\text{O}$  D.  $\text{C}_4\text{H}_{10} / \text{CO}_2$  E.  $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$  2. Given The Following Equation:  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$  A. How Many Moles Of  $\text{O}_2$  Can Be Produced By ...File Size: 80KBPage Count: 2 Apr 2th, 2024

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Name: KEY 33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1\text{mol}$

$122.55\text{g} \times 2\text{O}_2 \times 32.00\text{g} \times 1\text{mol} = 3.11\text{g O}_2$

% Yield =  $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$  34.  $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Percent purity = 62.8%? pure  $\text{FeCO}_3$  Apr 2th, 2024

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44. What Number Of Moles Of  $O_2$  Is Needed To  
Produce 14.2 Grams Of  $P_4O_{10}$  From  $P$ ? (Molar Mass  $P_4O_{10} = 284$ ) (A) 0.0500 Mole (B) 0.0625 Mole (C)



0.125 Mole May 2th, 2024

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