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Compounds Are Dissolved In Water, The Following  
Equilibrium Is Established.  $V^{2+} + H_2O \rightleftharpoons V^{3+} + H_2O$   
 $V^{2+} + OH^-$  What Would Alter The Composition Of The  
Equilibrium Mixture In Favour Of The  $V^{2+}$  Ions? A  
Adding An Acid B Adding A Reagent That Selectively

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For The Lattice Energy Of AgBr(s). First Ionisation  
Energy Of Silver = +731kJmol<sup>-1</sup> Electron Affinity Of  
Bromine = 325kJmol<sup>-1</sup> Enthalpy Change Of  
Atomisation Of Silver = +285kJmol<sup>-1</sup> Enthalpy Change  
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