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Relationship, So 0.0322 Mol Of Aniline Produces 0.0322 Mol Of Acetanilide. The Amount Of Acetic Anhydride Does Not Play A Role, Since It Is In Excess. Then, $n(\text{N}) = 0.0322 \text{ mol aniline} \times \frac{1 \text{ mol Acetanilide}}{1 \text{ mol Aniline}} = 0.0322 \text{ mol}$. We Find The Molar Mass Of Acetanilide: $(12 \times 8) + (1 \times 9) + (14 \times 1) + (16 \times 1) + 135 = 179 \text{ g/mol}$.
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2. Determination Of Cu(II) Content In $\text{Cu}(\text{gly})_2 \cdot x\text{H}_2\text{O}$

(1) Weigh 1.0 1.2 G (precision Of ± 0.0002 G) Of $\text{Cu}(\text{gly})_2 \cdot x\text{H}_2\text{O}$ With A Dry 100 cm^3 Beaker As The Container. (2) Dissolve It With 40 cm^3 Of Water And 8 cm^3 Of H_2SO_4 (1.0 $\text{mol} \cdot \text{dm}^{-3}$). (3) Transfer The Above Solution Quantitatively To A 100 cm^3 Volumetric Feb 3th, 2024

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