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Was Previously Published A Feb 2th, 2024  
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 Name: \_\_\_KEY\_\_\_ 33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$   
 Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1\text{mol}$   
 $122.55\text{g} \times 2\text{O}_2 \times 32.00\text{g} \times 1\text{mol} = 3.11\text{g O}_2$   
 $\% \text{Yield} = \frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$  34.  $4\text{FeCO}_3 + \text{O}_2$   
 $\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$   
 Percent purity = 62.8%? pure  $\text{FeCO}_3$  Mar 1th,  
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 Janssen ... In A Camping Stove? 2. How Many Moles Of

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Shown In Figure 12.2.1. We Can Use The Balanced  
Chemical Equation For The Reaction And Either The  
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To Produce 14.2 Grams Of P<sub>4</sub>O<sub>10</sub> From P? (Molar  
Mass P<sub>4</sub>O<sub>10</sub> = 284) (A) 0.0500 Mole (B) 0.0625 Mole  
(C) 0.125 Mole Jan 3th, 2024.  
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