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Conductivity Part 1: Electrolytes And Non-Electrolytes
3. In Group B, Do All Four Compounds Appear To Be Molecular, Ionic, Or Molecular Acids? Classify Each As A Strong Or Weak Electrolyte, And Arrange Them From The Strongest To The Weakest, Based On Conductivity Values. 4. Write An Equation For The Dissociation Of Each Of The Compound May 2th, 2024.
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TABLE OF CONJUGATE ACID-BASE PAIRS Acid Base Ka (25 C)
TABLE OF CONJUGATE ACID-BASE PAIRS Acid Base K A (25 OC)
 HClO_4 ClO_4^- - H_2SO_4 HSO_4^- - HCl Cl^- - HNO_3 NO_3^- - H_3O^+ + H_2O H_2CrO_4 HCrO_4^- - 1.8×10^{-1} $\text{H}_2\text{C}_2\text{O}_4$ (oxalic Acid) Jan 1th, 2024
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(aq) Acid Base Conjugate Conjugate Acid Base
Acid Base Conjugate Conjugate . Acid Base . 2) What Is The Strongest Base In The Following Reaction?
 $\text{HNO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{NO}_3^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$
 H_2O Is The Strongest Base. Strong Acids, Such As HNO_3 Have Weak Conjugate Bases, So NO_3^- is A Weak Base. H_2O And Apr 3th, 2024

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