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$N\text{O} + n\text{e}^- \rightleftharpoons \text{NRR}$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G^0 + RT \ln \frac{A_{\text{NRR}}}{A_{\text{NO}} A_{\text{H}^+}}$ (2.1.38) Which Since $\Delta G = -nF E_{\text{rxn}}$ And $A_{\text{H}^+} = A_{\text{H}_2} = 1$ Becomes, $E = E^0 + \frac{RT}{nF} \ln \frac{A_{\text{NRR}}}{A_{\text{NO}}}$. (2.1.40) This Expression Is Called The

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