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123.88 G/mol 70.90 G/mol 137.32 ... May 11th, 2024

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OBJECTIVES: • Memorize The Values For STP. • Memorize And Be Able To Apply The Gas Laws: Boyle's, Charles, Dalton's Law Of Partial Pressure, Combined Gas Law, Gay-Lussac's, And Graham's. • Be Able To Use M Jan 9th, 2024

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Total Moles Of Gas =  $1.3 \text{ Mol} + 1.5 \text{ Mol} = 2.8 \text{ Mol}$  .  $P = (2.8 \text{ mol}) \times (0.0821 \text{ L-atm Mol-K}) \times (273 \text{K}) / 10 \text{ L} = 6.3 \text{ Atm}$  Or  $4.8 \times 10^3 \text{ Torr}$ . Title: WORKSHEET 16 Name \_\_\_\_\_  
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**For Gas Types: Natural Gas E, LL And Liquid Petroleum Gas B/P**

- The Burner Must Not Be Used Outside Of Its Capacity Range (see Capacity Graphs, Ch. 8.2).
- The Gas Supply Pressure Must Not Exceed The Gas Pressure Given On The Burner Plate.
- The Burner Is Not Preset. Burner Type Forced Draught Burners With Two Sta May 7th, 2024

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Worksheet 7 - Ideal Gas Law I. Ideal Gas Law The Findings Of 19th Century Chemists And Physicists, Among Them Avogadro, Gay-Lussac, Boyle And Charles, Are Summarized In The Ideal Gas Law:  $PV = NRT$   $P =$  Pressure  $V =$  Volume  $N =$  Moles Of Gas,  $R =$  Universal Gas Constant  $T =$  Temperature. The Value Of  $R$  Varies With The Jan 7th, 2024

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