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42 Chapter 2 Environmental Chemistry Section 2.2 Stoichiometry 43 We Already Know That There Are 44 Grams Per Mole Of CO₂ So We Do Not Need To Recalculate That. Two Moles Of Butane (2 Mol X 58 G/mol = 116 G) Yields 8 Moles Of Carbon Dioxide (8 Mol X 44 G/mol = 352 G CO₂) , So We Can Set Up The Following Proportion: Jun 1th, 2024

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Chemistry 11 Stoichiometry Review Worksheet Name:___KEY___ 33. KClO₄ → 2O₂

$2\text{KClO}_4 \rightarrow 2\text{KCl} + 2\text{O}_2$ Actual yield = 1.45g Theoretical yield = 5.95g
 $\text{KClO}_4 \times 1\text{mol} = 122.55\text{g}$
 $\text{O}_2 \times 2 = 32.00\text{g}$
 $1\text{mol} = 3.11\text{g}$
 $\text{O}_2 \text{ \%Yield} = \frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$
 $3\text{Fe} + 4\text{CO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$ Percent purity = 62.8%
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The Compound Para-aminobenzoic Acid (you May Have Seen It Listed As PABA On Your Bottle Of Sunscreen) Is Composed Of Carbon (61.31%), Hydrogen (5.14%), Nitrogen (10.21%), And Oxygen (23.33%). Find The ... - In Other Words It's The Amount Of Product Possible From Stoichiometry. The "perfect
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