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 Nitrogen (N₂) 100 150 200 250 300 350 450 550 700 1,000 1,100 1,200 1,300
 Oxygen (Apr 7th, 2024Introduction To Combustion Turns 3rd SolutionsIntroduction
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 The Sum Of The Number Of Moles Of All The Components Of The Mixture: $N = N_1 +$

$N_2 + \dots$ It follows that the sum of all the mole fractions of the mixture must also equal 1. $X_1 + X_2 + \dots = 1$ The mass of component i in a mixture is the product of the number of moles of i and its molecular weight, M_i . The mass of the mixture is therefore the sum, $M = N_1M_1 + N_2M_2 + \dots$ Apr 5th, 2024.

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