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$2 \rightarrow \text{CaSiO}_3$  B.  $3\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow 2\text{Fe}_3\text{O}_4 + \text{CO}_2$  C.  $\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3$  D.  $\text{FeO} \rightarrow \text{Fe} + \frac{1}{2}\text{O}_2$  A. A B. D C. C And D D. A And D Answer: C Solution: In Metallurgy of Iron, CaO Is Used As Flux Which Is Used To Remove The Impurities Of  $\text{SiO}_2$ ,  $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$ . Also Here  $\text{Fe}_2\text{O}_3$  Is Reduced By CO To  $\text{Fe}_3\text{O}_4$  Which Is Further Reduced To ... Apr 5th, 2024

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