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 $2 \rightarrow \text{CaSiO}_3$   
B.  $3\text{Fe}_2\text{O}_3 + \text{CO} \rightarrow 2\text{Fe}_3\text{O}_4 + \text{CO}_2$   
C.  $\text{FeO} + \text{SiO}_2 \rightarrow \text{FeSiO}_3$   
D.  $\text{FeO} \rightarrow \text{Fe} + \frac{1}{2}\text{O}_2$   
A. A B. D C. C And D D. A And D  
Answer: C  
Solution: In Metallurgy Of Iron, CaO Is Used As Flux Which Is Used To Remove The Impurities Of SiO<sub>2</sub>,  
 $\text{CaO} + \text{SiO}_2 \rightarrow \text{CaSiO}_3$ . Also Here Fe<sub>2</sub>O<sub>3</sub> Is Reduced By CO To Fe<sub>3</sub>O<sub>4</sub> Which Is Further Reduced To ...  
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