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CHEN E4010 Mathematical Methods In Chemical EngineeringFrancis B. Hildebrand, Advanced Calculus For Applications, 2nd Edition, Prentice-Hall (1976). ISBN-13: 978-0130111890, ISBN-10: 0130111899 B. Materials(required): 1. CHEN E4010

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Introduction To Chemical Engineering: Chemical Reaction ...1 Chemical Reactions  
1.1 Rate Of Reaction And Dependence On Temperature We Will Once Again Look At  
The Formation Of Ammonia ( $\text{NH}_3$ ) From Nitrogen And Hydrogen (see Section  
Chemical Equilibrium Of The Thermodynamics Chapter). This Reaction Follows The  
Equation:  $\text{N}_2 + 3\text{H}_2 \rightleftharpoons 2\text{NH}_3$  (1)  $\Delta H^\circ = -92 \text{ kJ mol}^{-1}$   $\Delta S^\circ = -192 \text{ J mol}^{-1} \text{ K}^{-1}$  May 2th,  
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Chemical Engineering, CM3510: Chemical Reaction ...Grading: Exam 1 15%  
Exam 2 15% Exam 3 15% Final Exam 20% Research Paper 10% Homework 20%  
Class Participation 5% Last Year, The Following Grading Scale Was Used:  $A \geq 92$   $A/B$   
 $\geq 89$   $B \geq 82$   $B/C \geq 80$   $C \geq 71$   $C/D \geq 69$   $D \geq 61$   $F \leq 60$  If Necessary, The Grading  
Scale May Be Curved Downward. Apr 11th, 2024

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10.34: Numerical Methods Applied To Chemical Engineering

$\sum_{j=1}^K A_{jK} \frac{dC_j}{dt} = \sum_{j=1}^K F_{jK} - V \frac{dC_K}{dt} + R_K$

The sum of fluxes through each face of the volume  $\sum_{j=1}^K A_{jK} \frac{dC_j}{dt} = \sum_{j=1}^K F_{jK} - V \frac{dC_K}{dt} + R_K$ . We want to solve for  $C_K(t)$  by approximating the reaction and flux terms. Let's construct low order approximations physically.

$\frac{dC_K}{dt} = \frac{1}{V} \left( \sum_{j=1}^K F_{jK} - R_K \right)$

ACC IN/OUT GEN/CON \*  $\frac{dC_K}{dt} = \frac{1}{V} \left( \sum_{j=1}^K F_{jK} - R_K \right)$

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Form 1.2 Q.1 1.4 Q.1, 2 Miscellaneous Q.5 Write The Truth Values Of Statements 1.2 Q.2 1.4 Q.3, 5 1.6 Q.1 Mar 20th, 2024.

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