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$2\text{KCl}$  Actual yield = 1.45g Theoretical yield = 5.95g  $\text{KClO}_4 \times 1\text{mol} \times 122.55\text{g} \times 2\text{O}_2 \times 1\text{KClO}_4 \times 32.00\text{g} \times 1\text{mol} = 3.11\text{g}$   $\text{O}_2$  %Yield =  $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$   $34.4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$  Percent purity = 62.8%? g pure  $\text{FeCO}_3$  2th, 2024

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