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The Periodic Table And Periodic LawSection 6.2 Classification Of The Elements Key Concepts • The Periodic Table Has Four

Blocks (s, p, d, f). • Elements within a group have similar chemical properties. • The group number for elements in groups 1 and 2 equals the number of valence electrons. • Trends in the periodic table: In each group, the elements have similar properties, with a gradual trend in properties down the group. This is because they have the same number of electrons in the outer shells. Elements on the left side of the periodic table are metals and those on the right side are non-metals. Unit 3.2: The Periodic Table and Periodic Trends Notes. The organization of the periodic table. Dmitri Mendeleev was the first to organize the elements by their periodic properties. In 1871 he arranged the elements in vertical columns by their atomic mass and found he could get horizontal groups of 3 ...

Chemistry: Periodic Table & Periodic Trends Notes History of the periodic table Dimitri Mendeleev, a Russian chemist, is credited with creating the periodic table of elements in 1869. His breakthrough came when he arranged the elements in rows and columns based on their atomic mass and chemical properties.

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