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Oxidizing Agent And The Reducing Agent Also. Worksheet 7 Oxidationreduction Reactions Oxidation Number Rules. Choose A Method And Complete Questions 12 13. A Change In Oxidation Number B.  $\text{Si} + 2\text{F}_2 \rightarrow \text{SiF}_4$  8. Determine What Is Oxidized And What Is Reduced In Each Reaction.  $2\text{CS} + \text{Br}_2 \rightarrow 2\text{CSBr}$  4. Worksheet 25 Oxidationred Feb 9th, 2024

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State The Oxidation Number Of Each Of The Elements That Is Underlined. A)  $\text{NH}_3$  B)  $\text{H}_2\text{SO}_4$  C)  $\text{ZnCO}_3$  D)  $\text{Al}(\text{OH})_3$  E)  $\text{Na}_2\text{O}$  F)  $\text{Cl}_2$  G)  $\text{O}_2$  H)  $\text{Fe}_2\text{O}_3$  I)  $\text{H}_2\text{O}_2$  J)  $\text{H}_2\text{SO}_4$  K)  $\text{H}_2\text{SO}_4$  L)  $\text{H}_2\text{SO}_4$  M)  $\text{H}_2\text{SO}_4$  N)  $\text{H}_2\text{SO}_4$  O)  $\text{H}_2\text{SO}_4$  P)  $\text{H}_2\text{SO}_4$  Q)  $\text{H}_2\text{SO}_4$  R)  $\text{H}_2\text{SO}_4$  S)  $\text{H}_2\text{SO}_4$  T)  $\text{H}_2\text{SO}_4$  U)  $\text{H}_2\text{SO}_4$  V)  $\text{H}_2\text{SO}_4$  W)  $\text{H}_2\text{SO}_4$  X)  $\text{H}_2\text{SO}_4$  Y)  $\text{H}_2\text{SO}_4$  Z)  $\text{H}_2\text{SO}_4$  Balance The Redox

Equation Using The Half Reaction Method.  $\text{Al} + 3\text{Ag}^+ \rightarrow \text{Al}^{3+} + 3\text{Ag}$  18. Ci May 1th, 2024

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