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Review Answers Calcium $\text{C}_2\text{O}_2 + 3 \text{Cr}(3)$ When The
Potassium Chlorate (KClO_3) Is Heated, It Decomposes
To Form Gases KCl And Oxygen (O_2 .) $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$
C. Mar 1th, 2024Chapter 12 Stoichiometry
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11 Stoichiometry Review Worksheet Name: ____KEY____

33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g

Theory yield = $5.95\text{g KClO}_4 \times 1\text{mol} / 122.55\text{g} \times 2 \times 1\text{KClO}_4$

$4 \times 32.00\text{g} / 1\text{mol} = 3.11\text{g O}_2$ %Yield = $1.45\text{g} / 3.11\text{g}$

$\times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$

Percent purity = 62.8%? pure FeCO_3 Apr 1th, 2024.

Stoichiometry Worksheet 2 Percent Yield

Answers Glencoe Study Guide Answer Key , Pearson

Chemistry Workbook Answers , Nakamichi 582 User

Manual , Walmart Application Answers , Acls

Certification Study Guide , Deltek Electronic Expense

Training Manual v3 , Compaq Evo N620c Pc Notebook

Manual , American English File 3 Feb 8th,

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Certification User Guide Support, Beating The Street,

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Can Use The Balanced Chemical Equation For The
Reaction And Either The Masses Of Solid Reactants
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Sodium Sulfate Will Be Formed If You Start With 200.0.
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