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### **Structural Formulas (Section 1-10) Full Structural ...**

4. Draw The Full Structure, Given The Condensed Structure. (Note: CH<sub>3</sub>CH<sub>2</sub>OH (CH<sub>3</sub>)<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub> CH<sub>2</sub>CHCl CH<sub>3</sub>CHO CH<sub>3</sub>CO<sub>2</sub>H 5. Fill In The Full Structure, Including Attached Hydrogens And Attached Lone Pairs, For The Following Line-angle Structures. If Given A Condensed Structure, Convert It To A Line-angle. OH O CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub> CH<sub>3</sub>CO<sub>2</sub>H 4th, 2024

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$\cot x = -\csc^2 x$  (11)  $D_x \sec x = \sec x \tan x$  (12)  $D_x \csc x = -\csc x \cot x$  (13)  $D_x e^x = e^x$  (14)  $D_x \ln a = \frac{1}{a}$  (15)  $D_x \ln|x| = \frac{1}{x}$  (16)  $D_x \sin^{-1} x = \frac{1}{\sqrt{1-x^2}}$  (17)  $D_x \cos^{-1} x = \frac{-1}{\sqrt{1-x^2}}$  (18)  $D_x \tan^{-1} x = \frac{1}{x^2+1}$  (19)  $D_x \cot^{-1} x = \frac{-1}{x^2+1}$  (20) D 3th, 2024

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### **Draw The Correct Structural Formulas**

For Example  $CH_3COOH$  Or  $CH_2$  Instead Of Showing All These Bonds, Draw A Correct Lewis Structure For  $CH_3COOH$  See 35 The Two  $CH_3$  S Go At The Ends Vertically In The Middle Is The Linear  $HCOH$  With A Double Bonded O At The Second Carbon 36 Draw 1th, 2024

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### **2.8 Formulas, Names And Masses Structural Formula Of ...**

HCl As Ionic Or Covalent And Identify The Ions If Present. Problem: • Name The Compounds: PF<sub>5</sub>, CrF<sub>3</sub>, N<sub>2</sub>O<sub>3</sub>, Fe<sub>2</sub>O<sub>3</sub> - Molecular →phosphorus Pentafluoride - Ionic, F<sup>-</sup>, Cr<sup>3+</sup> →chromium(III) Fluoride - Molecular →dinitrogen Trioxide - Ionic, O<sub>2</sub><sup>-</sup>, Fe<sup>3+</sup> →iron(III 2th, 2024

### **Structural Formulas And Isomers**

2. The Compound C<sub>4</sub>H<sub>9</sub>OH Is An Isomer Of (1) C<sub>3</sub>H<sub>7</sub>COCH<sub>3</sub> (2) CH<sub>3</sub>COOC<sub>2</sub>H<sub>5</sub> (3) C<sub>2</sub>H<sub>5</sub>OC<sub>2</sub>H<sub>5</sub> (4) CH<sub>3</sub>COOH 3. If A Compound Has A Molecular Formula Of CH<sub>2</sub>O<sub>2</sub>, Then Its Structural Formula Must Be (1) (3) HO C O HHC O H O (2) (4) C O O H H HC O H O 4. The Structural Formulas HC CCCH H HHH H CCC H H C H H And H Represent Molecules Which Both Are (1) Halogen ...File Size: 128KBPage Count: 2 2th, 2024

### **2.1. How Do We Write Structural Formulas Of Alkanes?**

(b) 1,1-Dimethylcyclobutane Does Not Show Cis-trans Isomerism Because Only One Arrangement Is Possible For The Two Methyl Groups. Because Both Methyl Groups Are Bonded To The Same Carbon, They Must Be Trans To Each Other—one Above The Ring, The Other Below It. (c) 1,3-Dimethylcyclobutane Shows Cis-trans Isomerism. 1th, 2024

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