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$-(8.314 \text{ J K}^{-1} \text{ Mol}^{-1} \times 298 \text{ K}) \times \ln K_c$ So $K_c = 2.14 \times 10^{-90}$ Alternatively, K_c Can Be Obtained Directly From E O Using The Nernst Equation: $E_o = \frac{\ln C}{RT} \frac{K_c}{nF} = 0.54 \text{ V}$ $0.54 \text{ V} = \frac{11}{1} \frac{1}{C} (8.314 \text{ J K}^{-1} \text{ Mol}^{-1} \times 298 \text{ K}) \ln (10^{-90} / 485 \text{ Cmol})$ K 6th, 2024

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